

coupling across, *PCCC*,⁸ *POCC*,⁹ *POCH*,¹⁰ *CCCH*,¹¹ and *HCCN*¹² and for four-bond coupling across *HC-COP*.¹³⁻¹⁶ Preliminary measurements¹⁷ indicate that such relationships may exist for coupling across *CCCC* and *MCCC* where M is either lead or mercury.

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Additions and Corrections

Stable Carbocations. CXXX. Carbon-13 Nuclear Magnetic Resonance Study of Halocarbenium Ions. Degree of Halogen "Back-Donation" and Relative Stability of Halocarbenium Ions [*J. Amer. Chem. Soc.*, **94**, 3551 (1972)]. By GEORGE A. OLAH,* Y. K. MO, and Y. HALPERN, Department of Chemistry, Case Western Reserve University, Cleveland, Ohio 44106.

The ¹³C nmr shift of the carbocation center in the dimethylfluorocarbenium ion [(CH₃)₂C⁺F] was reported, due to a computational error made using the indor method, as -142.7 (from CS₂). The correct value is -89.1 (as also shown by recent FT measurements).

Conformational Analysis. XXVI. Conformational Equilibria in 5,5-Disubstituted 1,3-Dioxanes [*J. Amer. Chem. Soc.*, **94**, 8072 (1972)]. By ERNEST L. ELIEL* and RUDYARD M. ENANOZA, Department of Chemistry, University of Notre Dame, Notre Dame, Indiana 46556].

Professor Anteunis has informed us that Table I in ref 26, part 2 does not correspond to our Table II inasmuch as it refers to the equilibrium B ⇌ A, not A ⇌ B in Scheme II. Therefore the Δ*G*^o values for the last five entries in our Table II taken from ref 26 should read -0.19, -0.03, -0.32, -0.38, -0.17 kcal/mol, and the signs of Δ*G*^o in Table II do agree with those in Table III, entries 2, 5, 6, and 7. Moreover, in the case of 5-methyl-5-phenyl-1,3-dioxane the nmr method of ref 26 does lead to a clear choice as to which isomer predominates, based on the chemical shift of the 5-methyl group below the coalescence temperature.

We have also found that the infrared data reported on page 8076, Scheme III were subject to a calibration error; the correct frequencies are **27**, 3435 cm⁻¹; **28**, 3450, 3434 cm⁻¹. We are not sure whether the NH absorption in **27** indicates intramolecular H bonding;

rather the double absorption in **28** may be due to cis-trans isomerism of the amide. Configurational assignments are not affected by this uncertainty since they rest on the dipole determination of **19**. We thank Dr. W. F. Bailey for experimental assistance and Dr. Jan Stanek (Prague, Czechoslovakia) for helpful correspondence.

Excited State Carbonyl Species from the Thermal Decomposition of 3,3-Dibenzyl-1,2-dioxetane [*J. Amer. Chem. Soc.*, **94**, 9277 (1972)]. By WILLIAM H. RICHARDSON,* FREDERICK C. MONTGOMERY, and MARY B. YELVINGTON, Department of Chemistry, California State University, San Diego, San Diego, California 92115.

Figure 1, line 4 should read "bibenzyl" rather than "dibenzyl ketone."

Hydrolysis of Imidate Esters Derived from Weakly Basic Amines. Influences of Structure and pH on the Partitioning of Tetrahedral Intermediates [*J. Amer. Chem. Soc.*, **95**, 1253 (1973)]. By TADASHI OKUYAMA, TERRY C. PLETCHER, DAVID J. SAHN, and GASTON L. SCHMIR,* Department of Molecular Biophysics and Biochemistry, Yale University School of Medicine, New Haven, Connecticut 06510.

In Table I, the heading of column 5 should read 10⁻⁷*k*₇; the heading of column 7 should read 10⁻⁴·(*k*₃/*K*_a). The uppermost line in Figure 6 should be labeled 10⁻⁵*k*₇.

Stable Carbocations. CXLII. 1-Acenaphthenium Ions [*J. Amer. Chem. Soc.*, **95**, 3698 (1973)]. By GEORGE A. OLAH,* GAO LIANG, and PHILIP WESTERMAN, Department of Chemistry, Case Western Reserve University, Cleveland Ohio 44106.